

2,3-Dimethoxy-10-oxostrychnidinium 2-carboxy-4,5-dichlorobenzoate

Graham Smith,^{a*} Urs D. Wermuth^b and Jonathan M. White^c

^aSchool of Physical and Chemical Sciences, Queensland University of Technology, GPO Box 2434, Brisbane, Queensland 4001, Australia, ^bSchool of Biomolecular and Physical Sciences, Griffith University, Nathan, Queensland 4111, Australia, and

^cBIO-21 Molecular Science and Biotechnology, University of Melbourne, Parkville, Victoria 3052, Australia

Correspondence e-mail: g.smith@qut.edu.au

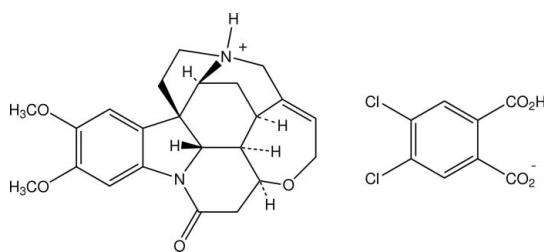
Received 28 September 2007; accepted 30 September 2007

Key indicators: single-crystal X-ray study; $T = 130\text{ K}$; mean $\sigma(\text{C}-\text{C}) = 0.003\text{ \AA}$; R factor = 0.036; wR factor = 0.087; data-to-parameter ratio = 12.4.

The structure of the title compound, $\text{C}_{23}\text{H}_{27}\text{N}_2\text{O}_4^+ \cdot \text{C}_8\text{H}_5\text{Cl}_2\text{O}_4^-$, a 1:1 proton-transfer compound of brucine with 4,5-dichlorophthalic acid, has been determined at 130 K. The brucinium cations and the hydrogen phthalate anions associate through single $\text{N}-\text{H} \cdots \text{O}_{\text{carboxylate}}$ hydrogen bonds [2.639 (3) Å], giving dimers which are extended *via* weak head-to-tail $\text{C}-\text{H} \cdots \text{O}_{\text{methoxy}}$ associations into chains forming down the 2_1 screw axis of the unit cell. The carboxyl proton of the anion gives a short intramolecular $\text{O}-\text{H} \cdots \text{O}_{\text{carboxylate}}$ hydrogen bond [2.441 (3) Å].

Related literature

Absolute configuration: (Peerdeman, 1956; Flack, 1983). Similar structures: (Oshikawa *et al.*, 2002; Smith *et al.*, 2005, 2006a,b; Bialońska & Ciunik, 2004a,b, 2006; Gould & Walkinshaw, 1984; Mallinson *et al.*, 2003; Bozkurt *et al.*, 2006).



Experimental

Crystal data

$\text{C}_{23}\text{H}_{27}\text{N}_2\text{O}_4^+ \cdot \text{C}_8\text{H}_5\text{Cl}_2\text{O}_4^-$
 $M_r = 629.47$
 Monoclinic, $P2_1$
 $a = 9.5085 (10)\text{ \AA}$
 $b = 12.9946 (13)\text{ \AA}$

$c = 11.5893 (12)\text{ \AA}$
 $\beta = 104.110 (2)^\circ$
 $V = 1388.8 (2)\text{ \AA}^3$
 $Z = 2$
 Mo $K\alpha$ radiation

$\mu = 0.29\text{ mm}^{-1}$
 $T = 130 (2)\text{ K}$

$0.45 \times 0.20 \times 0.15\text{ mm}$

Data collection

Bruker SMART CCD area-detector diffractometer
 Absorption correction: multi-scan (*SADABS*; Bruker, 1999)
 $T_{\min} = 0.86$, $T_{\max} = 0.96$

8761 measured reflections
 4923 independent reflections
 4747 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.024$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.036$
 $wR(F^2) = 0.087$
 $S = 1.04$
 4923 reflections
 396 parameters
 1 restraint

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\max} = 0.28\text{ e \AA}^{-3}$
 $\Delta\rho_{\min} = -0.27\text{ e \AA}^{-3}$
 Absolute structure: Flack (1983), 1604 Friedel pairs
 Flack parameter: 0.07 (5)

Table 1
 Hydrogen-bond geometry (Å, °).

$D-\text{H} \cdots A$	$D-\text{H}$	$\text{H} \cdots A$	$D \cdots A$	$D-\text{H} \cdots A$
N19—H19···O11A	0.90 (3)	1.75 (3)	2.639 (3)	170 (2)
O21A—H21A···O12A	0.84 (5)	1.61 (5)	2.441 (3)	170 (4)
C3A—H3A···O22A	0.93	2.31	2.667 (3)	103
C4—H4···O25	0.93	2.45	2.953 (3)	114
C6A—H6A···O11A	0.93	2.26	2.635 (3)	103
C16—H16···O12A	0.98	2.55	3.402 (3)	145
C17—H17B···O2 ⁱ	0.97	2.53	3.441 (3)	156
C18—H18B···O3 ⁱ	0.97	2.48	3.279 (3)	140
C22—H22···O22A ⁱⁱ	0.93	2.45	3.316 (3)	155

Symmetry codes: (i) $-x + 3, y + \frac{1}{2}, -z + 1$; (ii) $x, y + 1, z$.

Data collection: *SMART* (Bruker, 2000); cell refinement: *SMART*; data reduction: *SAINT* (Bruker, 1999); program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *PLATON* (Spek, 2003); software used to prepare material for publication: *PLATON*.

The authors acknowledge financial support from the School of Physical and Chemical Sciences (Queensland University of Technology), the School of Science, Griffith University and the School of Chemistry, the University of Melbourne.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: FL2168).

References

- Bialońska, A. & Ciunik, Z. (2004a). *Acta Cryst.* **C60**, o853–o855.
- Bialońska, A. & Ciunik, Z. (2004b). *CrystEngComm*, **6**, 276–279.
- Bialońska, A. & Ciunik, Z. (2006). *Acta Cryst.* **C62**, o450–o453.
- Bozkurt, E., Kartal, I., Odabaşoğlu, M. & Büyükgüngör, O. (2006). *Acta Cryst.* **E62**, o4258–o4260.
- Bruker (1999). *SADABS* (Version 2.03) and *SAINT* (Version 6.02). Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2000). *SMART*. Version 5.55. Bruker AXS Inc., Madison, Wisconsin, USA.
- Flack, H. D. (1983). *Acta Cryst.* **A39**, 876–881.
- Gould, R. O. & Walkinshaw, M. D. (1984). *J. Am. Chem. Soc.* **106**, 7840–7842.
- Mallinson, P. R., Smith, G. T., Wilson, C. C., Grech, E. & Wozniak, K. (2003). *J. Am. Chem. Soc.* **125**, 4259–4270.
- Oshikawa, T., Pochamroen, S., Takai, N., Ide, N., Takemoto, T. & Yamashita, M. (2002). *Heterocycl. Commun.* **8**, 271–274.

- Peerdeman, A. F. (1956). *Acta Cryst.* **9**, 824.
Sheldrick, G. M. (1997). *SHELXL97* and *SHELXS97*. University of Göttingen, Germany.
Smith, G., Wermuth, U. D., Healy, P. C. & White, J. M. (2006a). *Aust. J. Chem.* **59**, 320–328.
Smith, G., Wermuth, U. D., Young, D. J. & Healy, P. C. (2005). *Acta Cryst. E* **61**, o2008–o2011.
Smith, G., Wermuth, U. D., Young, D. J. & White, J. M. (2006b). *Acta Cryst. E* **62**, o1553–o1555.
Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.